

# Tris(phenanthroline- $\kappa^2N,N'$ )nickel(II) bis(malate) tetrahydrate

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**Key indicators**

Single-crystal X-ray study  
 $T = 273$  K  
 Mean  $\sigma(C-C) = 0.004$  Å  
 $R$  factor = 0.046  
 $wR$  factor = 0.141  
 Data-to-parameter ratio = 17.7

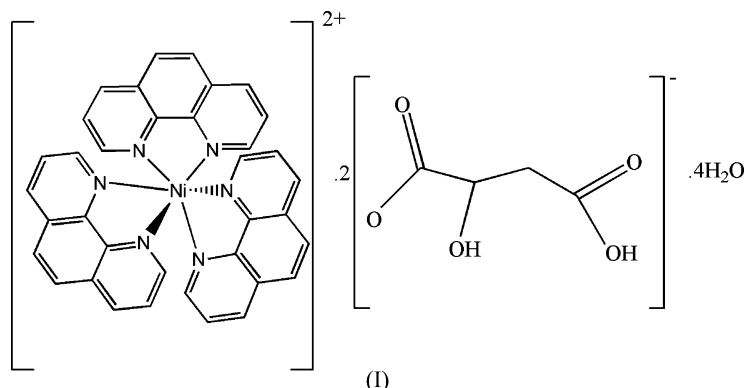
For details of how these key indicators were  
 automatically derived from the article, see  
<http://journals.iucr.org/e>.

The crystal structure of the title complex,  $[Ni(C_{12}H_8N_2)_3](C_4H_5O_5)_2 \cdot 4H_2O$ , consists of  $Ni^{II}$  complex cations, malate anions and solvent water molecules. The  $Ni^{II}$  ion is located on a twofold axis and chelated by three phenanthroline (phen) with a distorted octahedral geometry.

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**Comment**

$\pi$ - $\pi$  Stacking plays an important role in some biological processes. The title new  $Ni^{II}$  complex with 1,10-phenanthroline (phen) ligands, (I), is presented here.



The crystal structure of (I) consists of  $Ni^{II}$  complex cations, malate monoanions and solvent water molecules (Fig. 1). The  $Ni^{II}$  ion is located on a twofold axis and is chelated by three phen ligands in a distorted octahedral geometry (Table 1). The malate monoanions link the  $Ni^{II}$  complex cations *via* weak C—H...O hydrogen bonds (Table 2). Solvent water molecules and malate anions are linked together by O—H...O hydrogen bonds, which stabilize the crystal structure of (I).

**Experimental**

An aqueous solution (5 ml) of  $NiCl_2 \cdot 6H_2O$  (1 mmol) was mixed with an aqueous solution (5 ml) containing malic acid (2 mmol) and  $Na_2CO_3$  (2 mmol). The solution was refluxed for 30 min, then phen (2 mmol) was added. The mixture was refluxed for a further 30 min and filtered. Green single crystals of (I) were obtained from the filtrate after 6 d.

*Crystal data*

$[Cu(C_{12}H_8N_2)_3](C_4H_5O_5)_2 \cdot 4H_2O$	$Z = 4$
$M_r = 937.55$	$D_x = 1.468$ Mg m <sup>-3</sup>
Orthorhombic, <i>Pcca</i>	Mo $K\alpha$ radiation
$a = 19.7761$ (9) Å	$\mu = 0.54$ mm <sup>-1</sup>
$b = 14.4074$ (5) Å	$T = 273$ (2) K
$c = 14.8877$ (6) Å	Prism, green
$V = 4241.8$ (3) Å <sup>3</sup>	$0.40 \times 0.20 \times 0.16$ mm

## Data collection

Bruker SMART CCD area-detector diffractometer  
 $\varphi$  and  $\omega$  scans  
 Absorption correction: multi-scan (SADABS; Sheldrick, 1996)  
 $T_{\min} = 0.817$ ,  $T_{\max} = 0.918$

18394 measured reflections  
 5194 independent reflections  
 2617 reflections with  $I > 2I$   
 $R_{\text{int}} = 0.061$   
 $\theta_{\text{max}} = 28.1^\circ$

## Refinement

Refinement on  $F^2$   
 $R[F^2 > 2\sigma(F^2)] = 0.046$   
 $wR(F^2) = 0.141$   
 $S = 0.88$   
 5194 reflections  
 294 parameters

H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.078P)^2]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\text{max}} = 0.001$   
 $\Delta\rho_{\text{max}} = 0.32 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\text{min}} = -0.36 \text{ e } \text{\AA}^{-3}$

Table 1

Selected bond lengths (Å).

Ni1—N1	2.107 (2)	Ni1—N3	2.080 (2)
Ni1—N2	2.095 (2)		

Table 2

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1—H1A $\cdots$ O5 <sup>i</sup>	0.85	1.69	2.528 (3)	171
O3—H3A $\cdots$ O5	0.92	1.93	2.583 (3)	126
O6—H6A $\cdots$ O7	0.83	1.97	2.790 (9)	168
O6—H6B $\cdots$ O4	0.86	1.81	2.673 (6)	176
C8—H8 $\cdots$ O5 <sup>ii</sup>	0.93	2.49	3.415 (4)	177
C13—H13 $\cdots$ O1 <sup>iii</sup>	0.93	2.50	3.208 (4)	133
C14—H14 $\cdots$ O4 <sup>iv</sup>	0.93	2.58	3.301 (4)	135
C18—H18 $\cdots$ O2 <sup>v</sup>	0.93	2.43	3.327 (4)	162

Symmetry codes: (i)  $-x + \frac{3}{2}, y, z + \frac{1}{2}$ ; (ii)  $-x + 1, -y + 2, -z + 1$ ; (iii)  $-x + 1, y, -z + \frac{3}{2}$ ; (iv)  $x - \frac{1}{2}, y, -z + 1$ ; (v)  $x - \frac{1}{2}, -y + 1, -z + \frac{3}{2}$ .

H atoms on O1, O3 and O6 were located in a difference Fourier map and H atoms on O7 were placed in calculated positions ( $O-H = 0.85 \text{ \AA}$ ; Nardelli, 1999). H atoms on O atoms were refined as riding in as-found relative positions, with  $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$ . Other H atoms were placed in calculated positions, with  $C-H = 0.93$  (aromatic),  $0.97$  (methylene) or  $0.98 \text{ \AA}$  (methine), and refined in riding mode, with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ .

Data collection: SMART (Bruker, 1998); cell refinement: SAINT (Bruker, 1999); data reduction: SAINT; program(s) used to solve

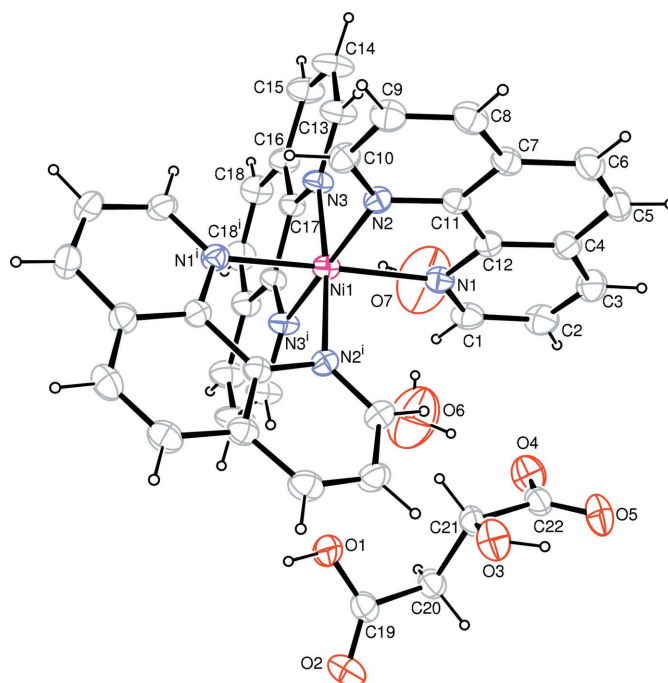


Figure 1

The molecular structure of (I), shown with 30% probability displacement ellipsoids (arbitrary spheres for H atoms) [symmetry code: (i)  $1 - x, y, \frac{3}{2} - z$ ].

structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: SHELXTL (Bruker, 1998); software used to prepare material for publication: SHELXTL.

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## References

- Bruker (1998). SMART (Version 5.0) and SHELXTL (Version 6.12). Bruker AXS Inc, Madison, Wisconsin, USA.  
 Bruker (1999). SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.  
 Nardelli, M. (1999). *J. Appl. Cryst.* **32**, 563–571.  
 Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.  
 Sheldrick, G. M. (1997). SHELXS97 and SHELXL97. University of Göttingen, Germany.